

Continue



11th grade review.

1. Solve the equation. Leave your answer as a simplest form of fraction.

$\frac{5}{2} = \frac{n}{6}$

Answer: \_\_\_\_\_

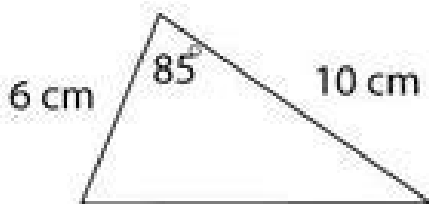
2.  $\frac{x}{5} = \frac{1}{4}$

Answer: \_\_\_\_\_

3.  $\frac{y}{5+1} = \frac{10}{3}$

Answer: \_\_\_\_\_

4. Find the area of the figure. Round to the nearest tenth.



Answer: \_\_\_\_\_

Boysen Cooking vs 3 Boil Questions

COOKING VS A BOIL

If you are bringing a big pot of cold water to a boil to cook some potatoes, to do it using the least amount of energy you should:  
a) turn the heat on full force, less chance for wasted heat.  
b) put the heat on only low if possible could save half of the time meaning it's less work!  
c) put the heat at some medium value

BOILING

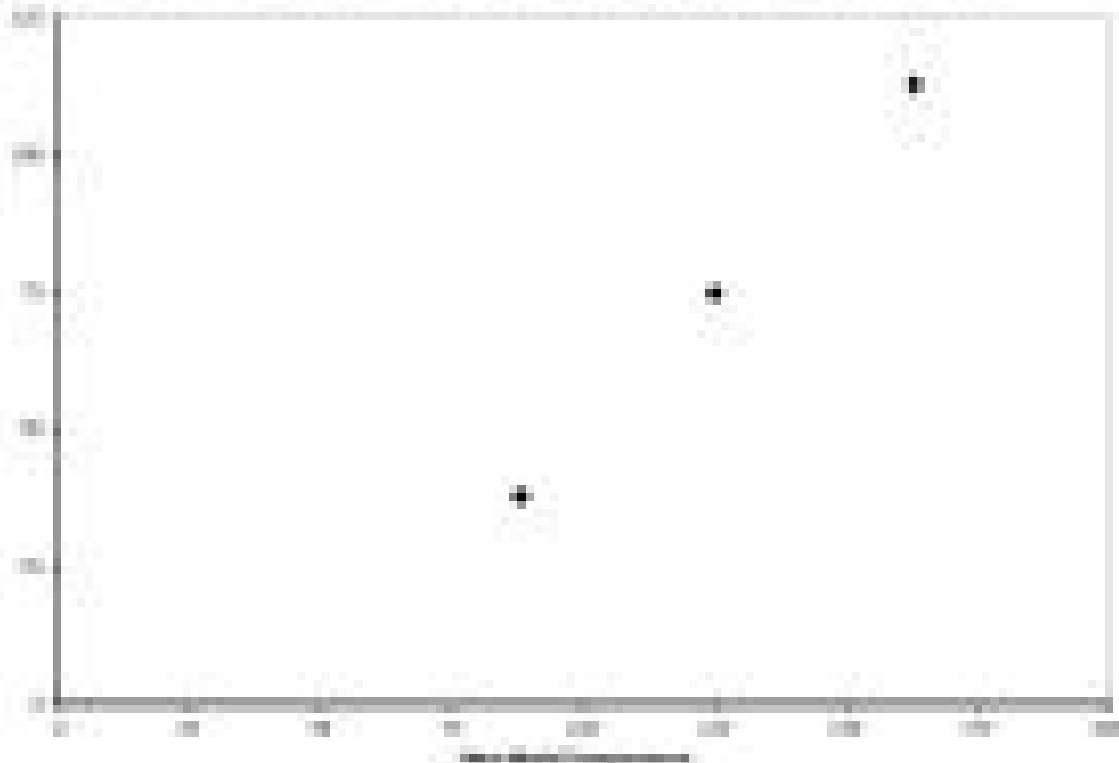
If the water is now boiling. To cook the potatoes using the least amount of energy you should  
a) keep the heat on full force, water energy just goes out as steam, doesn't make it faster (but putting a cover on would)  
b) turn the heat way down so the water just barely keeps boiling. (This is all that is needed to keep the water at boiling temp)  
c) keep the heat at some medium value

DEFLATION

If the volume occupied by some air is decreasing, then the temperature of the air must be  
a) increasing  
b) decreasing  
c) you can't tell. The answer also depends on how the pressure is changing. A balloon shrinking means gas with decreasing temperature, but air being compressed in an engine increases temperature.

NEW WORLD, NEW DATA

If you were to wake up in a "new world". In the new world, you have some measurements of the pressure in a tank of gas at various temperatures. A graph of your data looks like this:



Approximately what temperature is absolute zero in the "new world"?

- a) Zero New World Degrees  
b) 25 New World Degrees  
c) 40 New World Degrees

3. A 5.00 g mass of a metal was heated to 100.0 °C and then plunged into 100.0 g of water at 24.0 °C. The temperature of the resulting mixture became 28.0 °C. Cs for water= 4.184 J/gC

- a) How many joules did the water absorb?  
b) How many joules did the metal lose?  
c) What is the specific heat of the metal?  
d) What is the heat capacity of the 5.00 g sample?

4. Calculate the number of joules of heat released when 72.5 grams of water at 95.0°C cools to 28.0°C, specific heat capacity of water= 4.184 J/g C.

5. What is the specific heat of aluminum if the temperature of a 28.4 g sample of aluminum is increased by 8.1 °C when 207 J of heat is added?

6. A 13.5 g sample of gold is heated, then placed in a calorimeter containing 60.0 g of water. Temperature of water increases from 19.00 °C to 20.00 °C. The specific heat of gold is 0.130 J/g °C. What was the initial temperature of the gold metal sample?

7. A 5.00 g sample of acetylene undergoes complete combustion in an oxygen bomb calorimeter. The calorimeter has a heat capacity of 6.49 kJ/C. The maximum temperature increase is recorded as 9.50 C. Find the molar heat of combustion for acetylene, C<sub>2</sub>H<sub>2</sub> (Ans: -3.21 × 10<sup>3</sup> kJ/mol)

8. Solve the following Hess's law problems.

a. Calculate the ΔH for the following reaction: <b>Sn + 2Cl<sub>2</sub> à SnCl<sub>4</sub></b> You are given these two equations: Sn + Cl <sub>2</sub> à SnCl <sub>2</sub> ΔH = -325 kJ SnCl <sub>2</sub> + Cl <sub>2</sub> à SnCl <sub>4</sub> ΔH = -186 kJ	b. Calculate the ΔH for the following reaction: <b>NO + O à NO<sub>2</sub></b> You are given these three equations: O <sub>2</sub> à 2O ΔH = +495 kJ 2O <sub>2</sub> à 3O <sub>2</sub> ΔH = -427 kJ NO + O <sub>2</sub> à NO <sub>2</sub> + O <sub>2</sub> ΔH = -199 kJ
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9. Use a standard enthalpies of formation table to determine the change in enthalpy for each of these reactions.

- a) NaOH(s) + HCl(g) → NaCl(s) + H<sub>2</sub>O(g)  
b) 2 CO(g) + O<sub>2</sub>(g) → 2 CO<sub>2</sub>(g)

10. What is the change in enthalpy when 9.75 g of aluminum reacts with excess ammonium nitrate (NH<sub>4</sub>NO<sub>3</sub>) according to the equation:





<b>Worksheet 1.6</b>	<b>Leader: Lea</b>
<b>Supplemental Instruction</b>	<b>Course: Chem 177</b>
<b>Iowa State University</b>	<b>Instructor: Dr. Schwes-Müller</b>
	<b>Date: 2/5/2013</b>

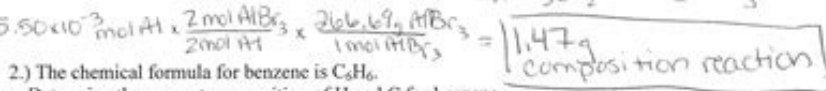
1.) A piece of aluminum foil 1.00 cm square and 0.550 mm thick is allowed to react with bromine to form aluminum bromide.  
a.) How many moles of aluminum were used? (density Al = 2.699 g/cm<sup>3</sup>)

$$1.00\text{ cm} \times 0.550\text{ cm} \times 0.055\text{ cm} = 0.03025\text{ cm}^3$$

$$\frac{0.03025\text{ cm}^3}{0.001\text{ cm}^3} \times 2.699\text{ g/cm}^3 = 0.0816\text{ g Al}$$

$$\frac{0.0816\text{ g Al}}{26.98\text{ g/mol Al}} = 0.003025\text{ mol Al}$$

b.) How many grams of aluminum bromide form, assuming the aluminum reacts completely? What type of reaction is this?



2.) The chemical formula for benzene is C<sub>6</sub>H<sub>6</sub>.

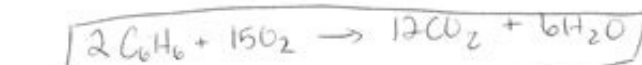
a. Determine the percent composition of H and C for benzene.

$$6(12.01) + 6(1.01) = 78.12$$

$$\text{C: } \frac{72.06}{78.12} \times 100 = 92.1\%$$

$$\text{H: } \frac{6.06}{78.12} \times 100 = 7.7\%$$

b. Write a balanced combustion reaction for benzene.



c. Assuming you have 364g of benzene, how many moles of O<sub>2</sub> will be required to completely combust this amount of benzene?

$$364\text{ g} \times \frac{1\text{ mol C}_6\text{H}_6}{78.12\text{ g C}_6\text{H}_6} \times \frac{15\text{ mol O}_2}{2\text{ mol C}_6\text{H}_6} = 35.0\text{ mol O}_2$$

d. How many grams of CO<sub>2</sub> will this produce? How many grams of H<sub>2</sub>O?

$$364\text{ g C}_6\text{H}_6 \times \frac{1\text{ mol C}_6\text{H}_6}{78.12\text{ g C}_6\text{H}_6} \times \frac{12\text{ mol CO}_2}{2\text{ mol C}_6\text{H}_6} \times \frac{44.01\text{ g CO}_2}{1\text{ mol CO}_2} = 1.23 \times 10^3\text{ g CO}_2$$

$$364\text{ g C}_6\text{H}_6 \times \frac{1\text{ mol C}_6\text{H}_6}{78.12\text{ g C}_6\text{H}_6} \times \frac{6\text{ mol H}_2\text{O}}{2\text{ mol C}_6\text{H}_6} \times \frac{18.02\text{ g H}_2\text{O}}{1\text{ mol H}_2\text{O}} = 2.52 \times 10^2\text{ g H}_2\text{O}$$

e. You perform the combustion, and 956g of CO<sub>2</sub> are produced. What is the percent yield?

$$\frac{\text{Actual}}{\text{Theoretical}} \times 100 = \frac{956\text{ g CO}_2}{1230\text{ g CO}_2} \times 100 = 77.7\%$$

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Name \_\_\_\_\_ Class \_\_\_\_\_ Date \_\_\_\_\_

#### Skills Worksheet

## Concept Review

### MATCHING

In the space provided, write the letter of the description that best matches the term or phrase.

- |  |  |
|--|--|
| _____ 1. practice of growing, breeding, and caring for plants and animals used for a variety of purposes     | <b>a.</b> loss of biodiversity         |
| _____ 2. study of how living things interact with each other and with their nonliving environments           | <b>b.</b> supply and demand            |
| _____ 3. conflict between short-term interests of individuals and long-term welfare of society               | <b>c.</b> “The Tragedy of the Commons” |
| _____ 4. declining number and variety of the species in an area  | <b>d.</b> agriculture                  |
| _____ 5. study of how humans interact with the environment   | <b>e.</b> developed nation             |
| _____ 6. law describing the relationship between an item’s availability and its value.                       | <b>f.</b> environmental science        |
| _____ 7. characterized by low population growth rate, high life expectancy, and diverse industrial economies | <b>g.</b> ecology                      |
| _____ 8. characterized by high population growth rate, low energy use, and very low personal wealth          | <b>h.</b> developing nation            |
| _____ 9. state in which a human population can survive indefinitely  | <b>i.</b> renewable resource           |
| _____ 10. natural material that can be replaced relatively quickly through natural processes                 | <b>j.</b> sustainability               |

### MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 11. Which of the following sciences contribute to the field of environmental science?
- |                                     |                            |
|-------------------------------------|----------------------------|
| <b>a.</b> physics and chemistry     | <b>c.</b> social sciences  |
| <b>b.</b> biology and earth science | <b>d.</b> all of the above |

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Holt Environmental Science **I** Science and the Environment

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